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Key indicators

Single-crystal X-ray study
 T = 297 K
 Mean $\sigma(C-C)$ = 0.005 Å
 R factor = 0.037
 wR factor = 0.082
 Data-to-parameter ratio = 13.1

For details of how these key indicators were automatically derived from the article, see <http://journals.iucr.org/e>.

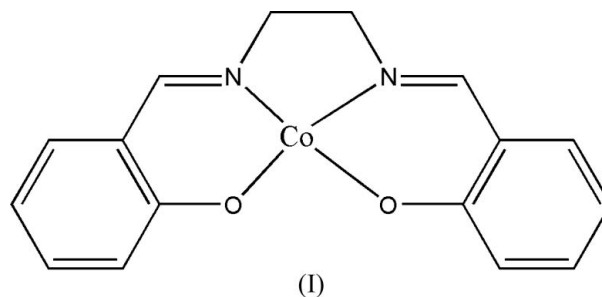
{2,2'-[Ethane-1,2-diylbis(nitrilomethylidene)]-diphenolato}cobalt(II)

The Co^{II} atom in the title compound, [Co(C₁₆H₁₄N₂O₂)], has a square-planar coordination.

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Comment

Cobalt(II) Schiff base complexes have been extensively utilized as catalysts for the oxidation of organic molecules (Mukaiyama & Yamada, 1995; Fiammengo *et al.*, 2002). We have previously reported some lanthanide complexes with Schiff base ligands (Yuan *et al.*, 2004). The title mononuclear cobalt(II) complex, (I), is a further contribution in this area. The Co^{II} atom is four-coordinate, chelated by two O and two N atoms in a square-planar geometry (Fig. 1).



Experimental

Red crystals of (I) were obtained by slow evaporation of a solution in ethyl acetate–methanol (1:1 v/v) of a mixture (10 ml) of *N,N'*-bis(-salicylidene)-1,2-ethylenediamine (0.048 g, 0.2 mmol) and cobalt diacetate tetrahydrate (0.050 g, 0.2 mmol).

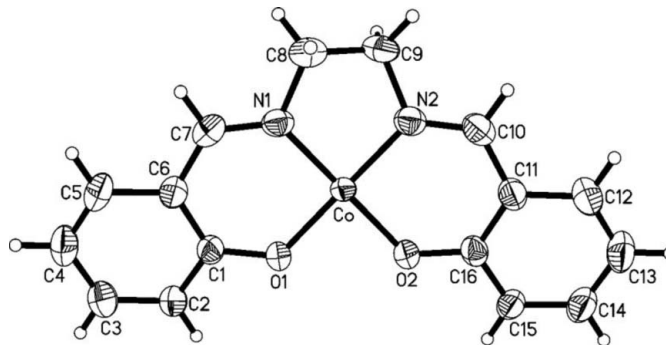


Figure 1
 The molecular structure of (I), showing the atom-numbering scheme. Displacement ellipsoids are drawn at the 30% probability level and H atoms are shown as small spheres of arbitrary radii.

Crystal data[Co(C₁₆H₁₄N₂O₂)] $M_r = 325.22$ Orthorhombic, *Pbca* $a = 7.471 (1) \text{ \AA}$ $b = 13.805 (2) \text{ \AA}$ $c = 26.096 (5) \text{ \AA}$ $V = 2691.5 (7) \text{ \AA}^3$ $Z = 8$ $D_x = 1.605 \text{ Mg m}^{-3}$ Mo $K\alpha$ radiation $\mu = 1.28 \text{ mm}^{-1}$ $T = 297 (2) \text{ K}$

Prism, red

 $0.58 \times 0.44 \times 0.12 \text{ mm}$ *Data collection*

Siemens P4 diffractometer

 ω scansAbsorption correction: ψ scan(North *et al.*, 1968) $T_{\min} = 0.514$, $T_{\max} = 0.858$

3059 measured reflections

2497 independent reflections

1605 reflections with $I > 2\sigma(I)$ $R_{\text{int}} = 0.013$ $\theta_{\max} = 25.5^\circ$

3 standard reflections

every 97 reflections

intensity decay: 2.0%

*Refinement*Refinement on F^2 $R[F^2 > 2\sigma(F^2)] = 0.037$ $wR(F^2) = 0.082$ $S = 0.97$

2497 reflections

191 parameters

H-atom parameters constrained

 $w = 1/[\sigma^2(F_o^2) + (0.035P)^2]$ where $P = (F_o^2 + 2F_c^2)/3$ $(\Delta/\sigma)_{\max} = 0.001$ $\Delta\rho_{\max} = 0.31 \text{ e \AA}^{-3}$ $\Delta\rho_{\min} = -0.31 \text{ e \AA}^{-3}$ Extinction correction: *SHELXL97*

Extinction coefficient: 0.0062 (3)

All H atoms were positioned geometrically and refined as riding, with C–H = 0.93 Å and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

Data collection: *XSCANS* (Siemens, 1994); cell refinement: *XSCANS*; data reduction: *SHELXTL* (Sheldrick, 1997b); program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997a); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997a); molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL*.

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References

- Fiammengo, R., Bruinink, C. M., Crego-Calama, M. & Reinhoudt, D. N. (2002). *J. Org. Chem.* **67**, 8552–8557.
- Mukaiyama, T. & Yamada, T. (1995). *Bull. Chem. Soc. Jpn.*, **68**, 17–35.
- North, A. C. T., Phillips, D. C. & Mathews, F. S. (1968). *Acta Cryst.* **A24**, 351–359.
- Sheldrick, G. M. (1997a). *SHELXS97* and *SHELXL97*. University of Göttingen, Germany.
- Sheldrick, G. M. (1997b). *SHELXTL*. Version 5.1. Bruker AXS Inc., Madison, Wisconsin, USA.
- Siemens (1994). *XSCANS*. Siemens Analytical X-ray Instruments Inc., Wisconsin, USA.
- Yuan, W.-B., Yan, L. & Yang, R.-D. (2004). *Chin. J. Appl. Chem.* **21**, 829–831.

Table 1

Selected bond lengths (Å).

Co–N2	1.838 (3)	Co–O1	1.845 (2)
Co–N1	1.844 (3)	Co–O2	1.850 (2)